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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle. C2H4(g) + H2(g) C2H6(g) 3 O 2 2CO2(g) + 3H2O - 1 4 1 1 k J m o l - 1 - 2 8 6 k J m o l - 1 - 1 5 6 0 k J m o l - 1 (hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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8 20 2 2 2 2 o correct reactants and products; [2] correct coefficients; M2 can only be scored if M1 correct. (b) Local pollutant: carbon monoxide/CO / volatile organics/VOCs / nitrogen oxide/NO / (unburnt) hydrocarbons; Do not accept methane/CH 4, ethane/C 2 H 6, propane/C 3 H 8 or butane/C 4 H 10. Global pollutant: nitrogen oxide/NO / carbon ...

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CHCHOH(l) 3O(g) 2CO(g) 3HO(l) -1367 kJ 3 2 + → + 2 2 2 H = mol ē Ö 1 -1 CH (l) 1 2O(g) 8CO(g) 9HO(l) 8 18 + → + 2 2 2 H =-5470 kJ mol ē Ö (i) Calculate the mass, in g, of ethanol and octane in 1.00 kg of the fuel mixture. [1].....

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possible liver damage; [2 max] (b) mild analgesics function by intercepting the pain stimulus at the source / interfere with the production of substances that cause pain/prostaglandins; strong analgesics work by bonding to receptor sites in the brain / prevent the transmission of pain impulses without depressing the central nervous system; [2]

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